



*Radiobioassay & Radiochemical Measurements Conference*



# Purity of DGA Normal for Po Separations

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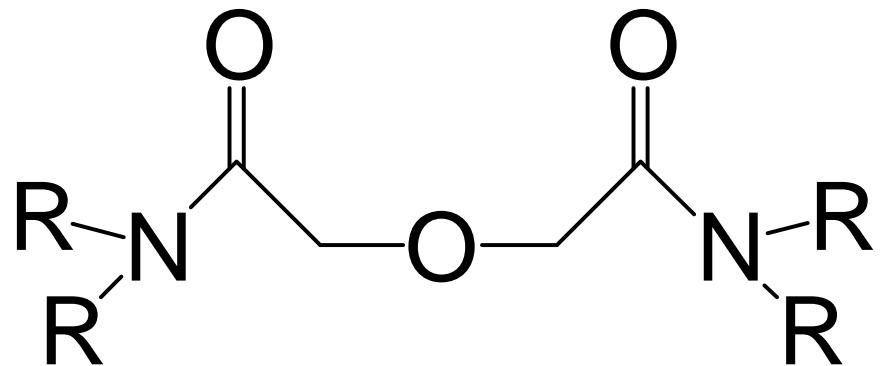
# DGA Applications

Rare Earth Separations<sup>1</sup>

Trivalent Actinides<sup>1</sup>

Removal of alpha  
emitters from Ra<sup>2</sup>

Polonium Separations<sup>3</sup>

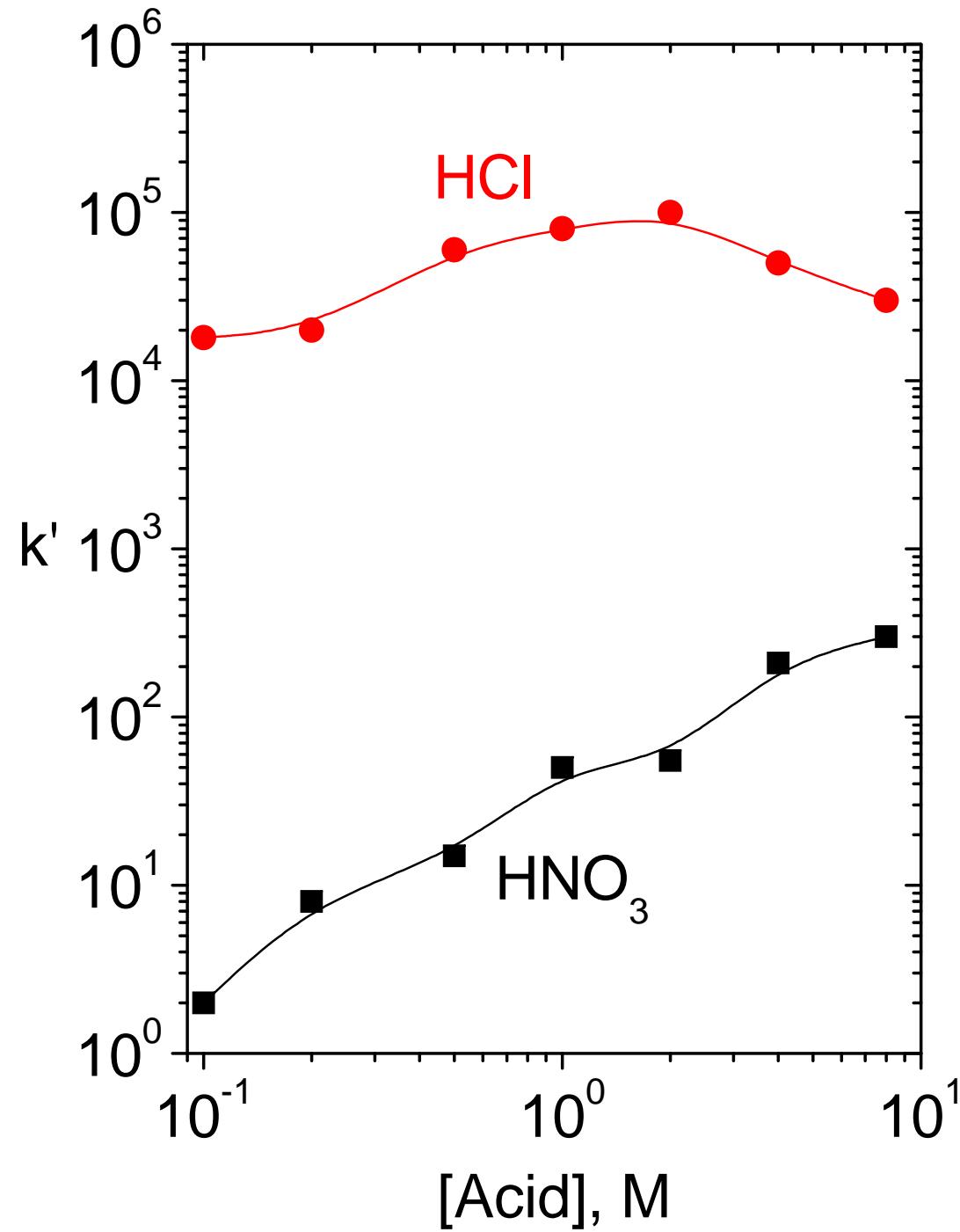


<sup>1</sup>Horwitz, McAlister, Bond, Barrans, Jr., *Solv. Extr. Ion Exch.*, 23, 319 (2005)

<sup>2</sup>Maxwell, Culligan, Hutchinson, Utsey, McAlister, *J. Radioan. Nucl. Chem.* 300(3), 1159 (2014)

<sup>3</sup>Maxwell, Culligan, Hutchinson, Utsey, McAlister, *J. Radioan. Nucl. Chem.* 298(3), 1977 (2014)

## $k'$ Po on DGA Resin



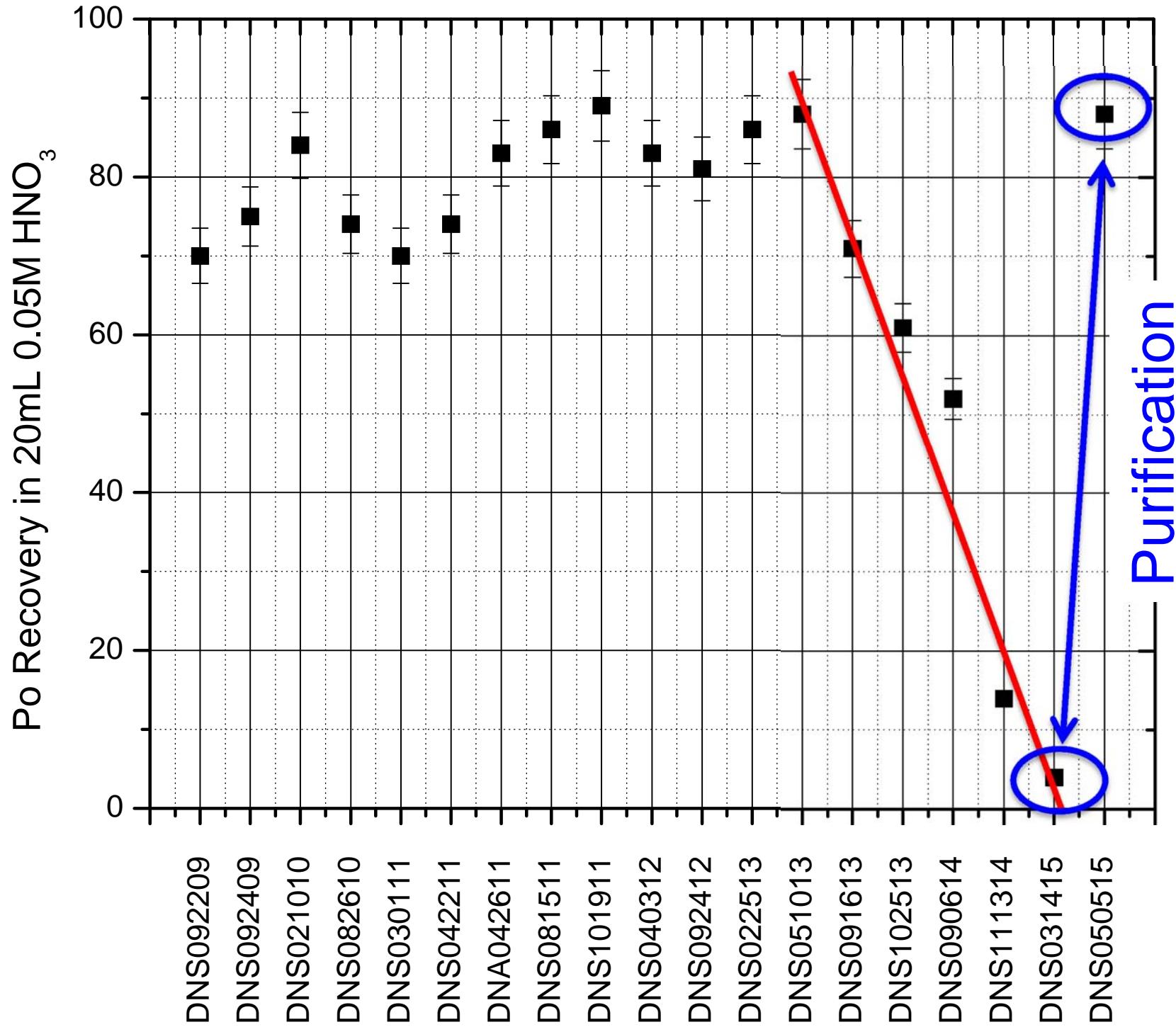
# DGA Quality Control

2009 and earlier: Elution of Eu-152 Tracer  
Periodic LC-MS

2010-2015: Separation of Am-241 and U-233

2015 and later: Add Po QC test

# Polonium Recovery on TODGA





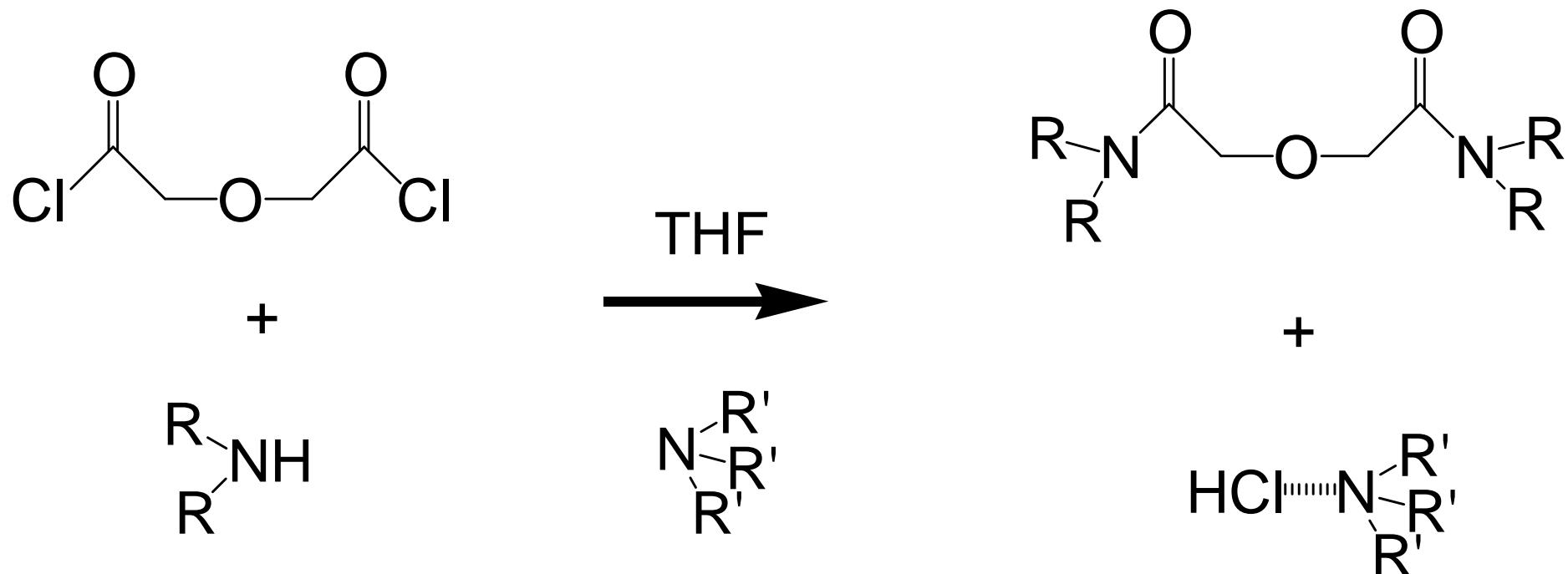
What Changed?

Can we purify?

How do we prevent?

Only Po???

# Synthesis of DGA from Diglycolyl chloride (DGC)



Horwitz, McAlister, Bond, Barrans, Jr., *Solv. Extr. Ion Exch.*, 23, 319 (2005)  
Sasaki, Sugo, Suzuki, Tachimori, *Solv. Extr. Ion Exch.*, 19, 91 (2001)

# Raw Materials

Purchased a large amount of DGC in 2011

- Synthesis Chemist noted darkening of material over time
- Dark material was not a problem in the past
- Dark material persists into final product
- Impurity is heat/pH sensitive

Supplier began packing Dioctylamine in plastic (was glass)

- Plastic containers frost over (amine + air/moisture)

No issue not observed with TEHDGA

- Is dioctylamine the source of the issue?

# Improved Synthesis and Purification

Distill Diglycolylchloride

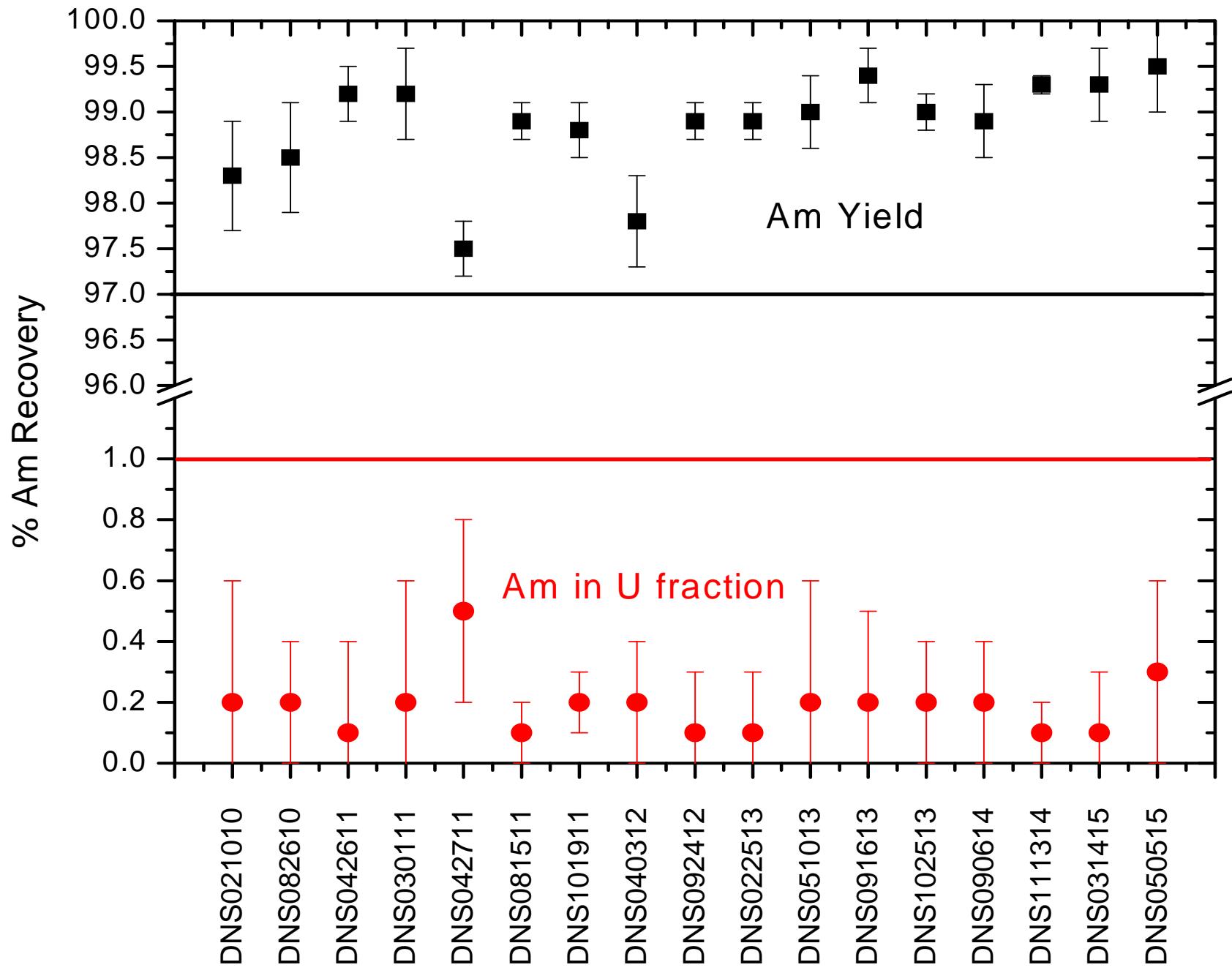
Distill Diethylamine

Dry Solvent

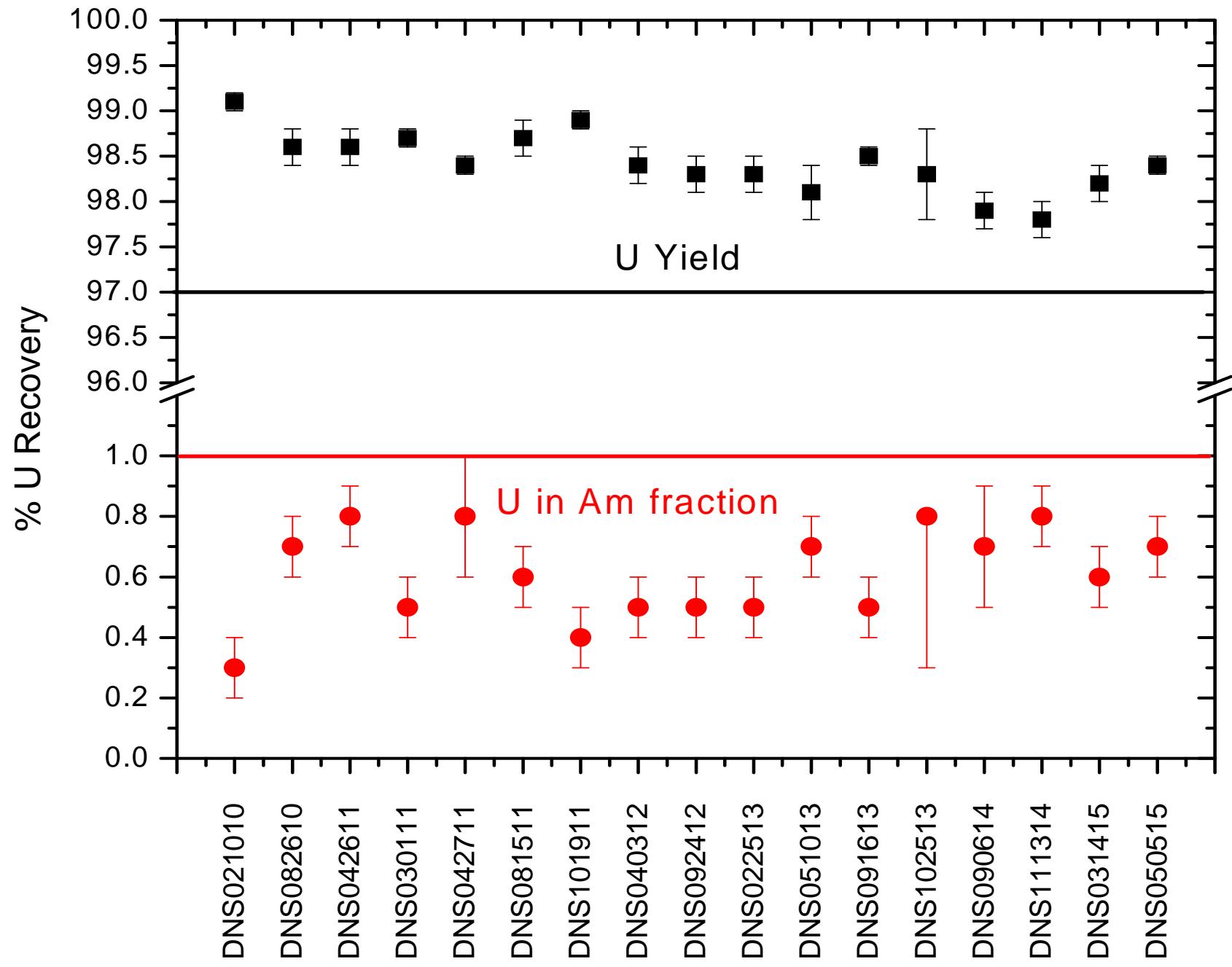
Test Po recovery.  
Additional purification as necessary.  
Full batch and standard QC.



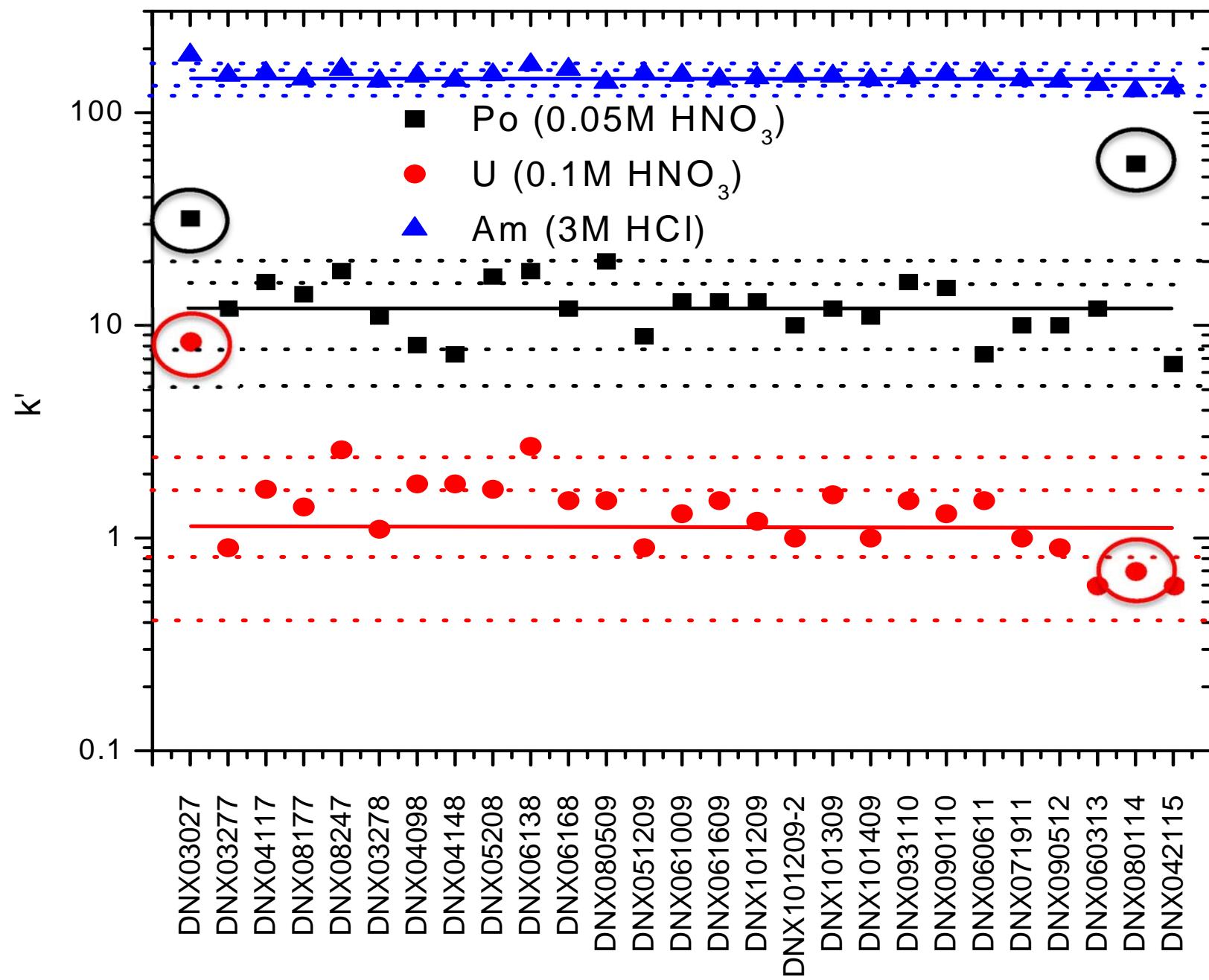
# Am in Standard DGA QC



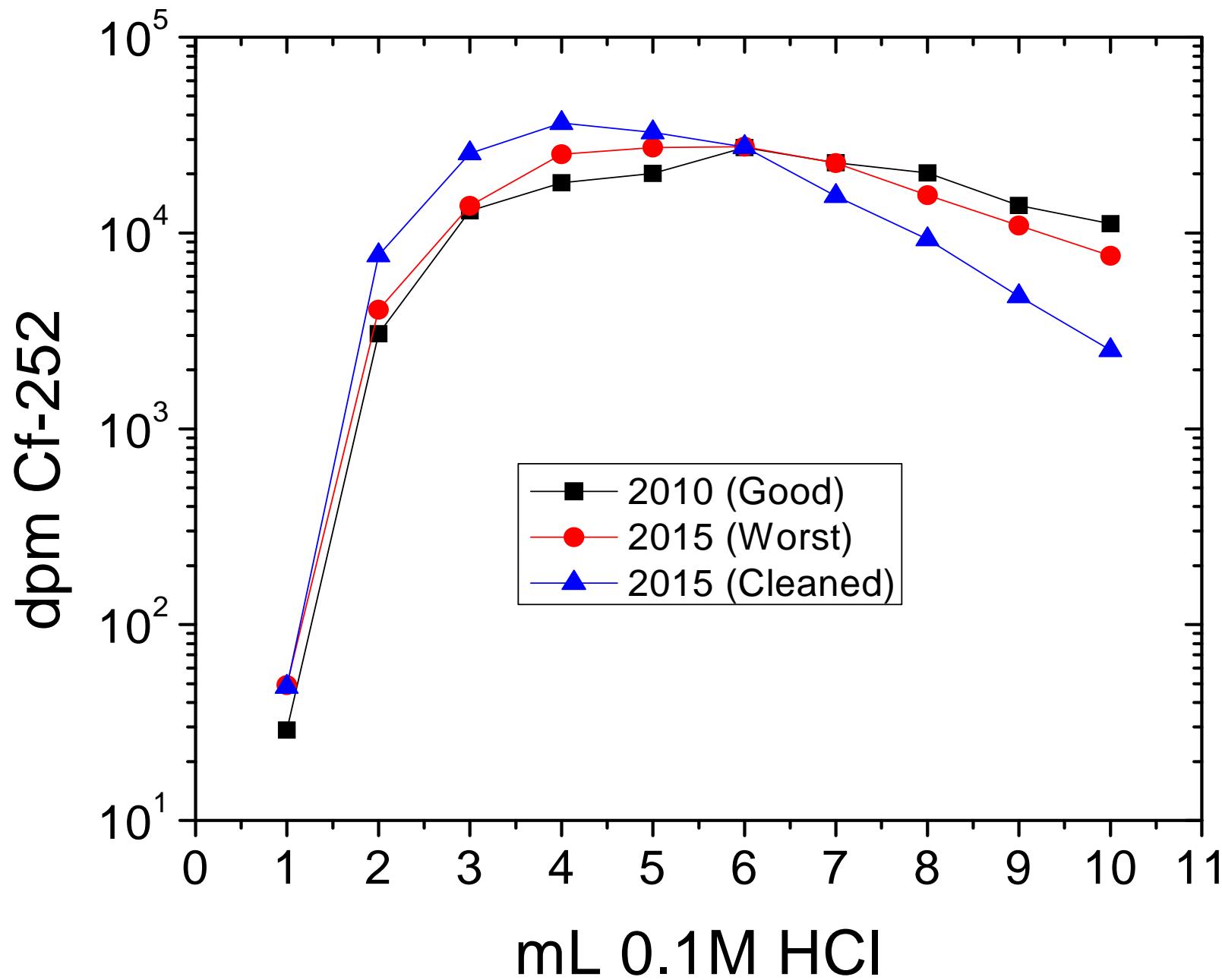
# U in Standard DGA QC

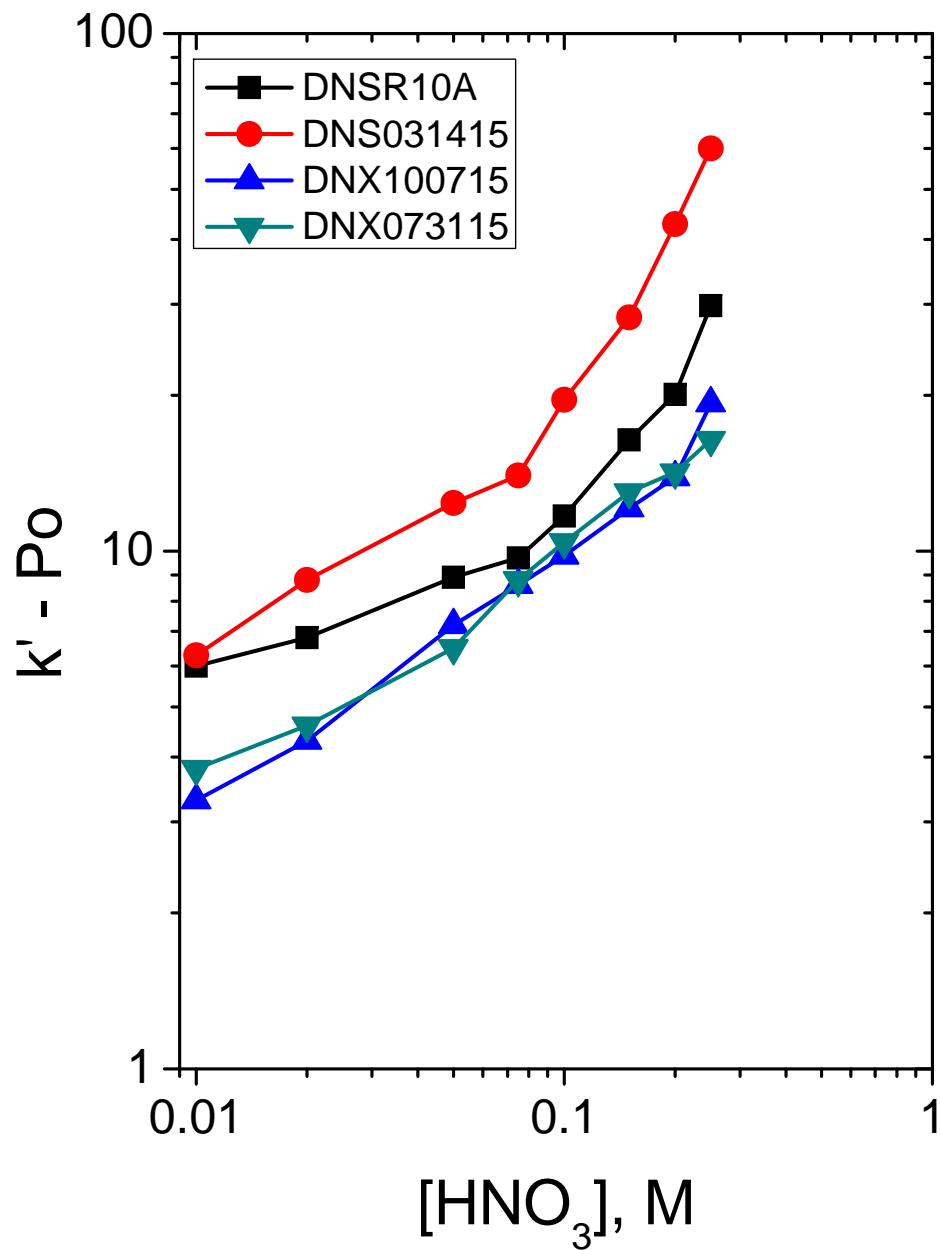


# $k'$ on DGA Resin Lots

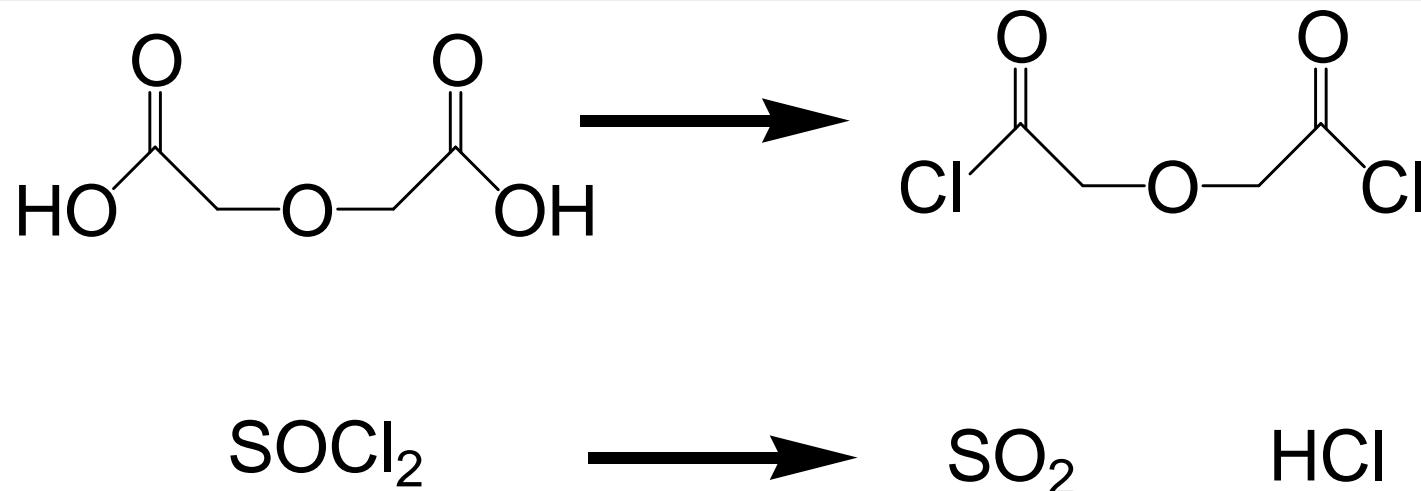


# Elution of Cf-252 from 2mL DGA, Normal Cartridges





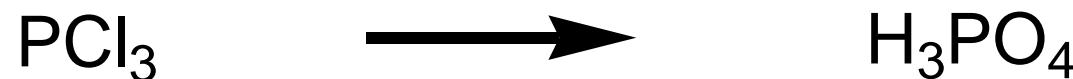
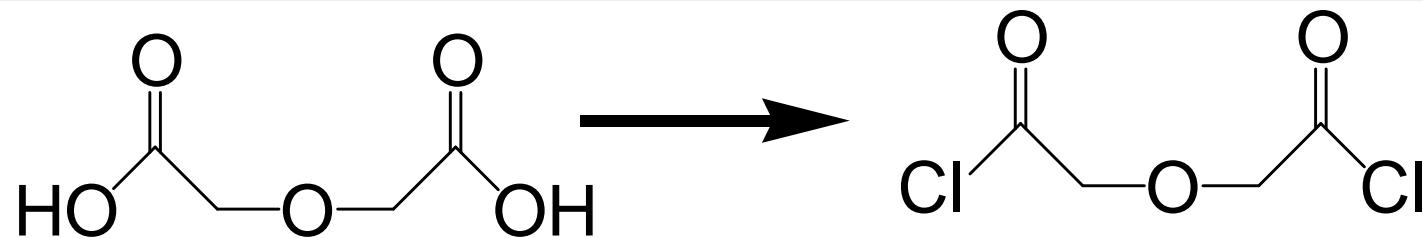
# Synthesis and Degradation of DGC



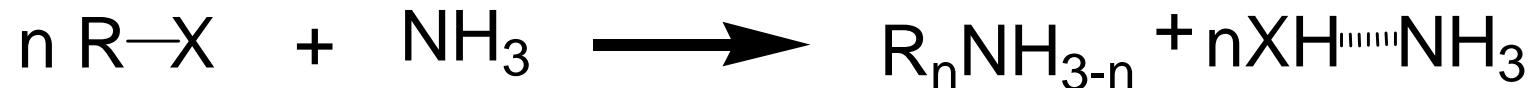
Aged samples of SO2Cl2 develop a yellow hue, possibly due to the formation of S2Cl2.

S2Cl2 has been used to introduce C-S bonds.

# Synthesis and Degradation of DGC



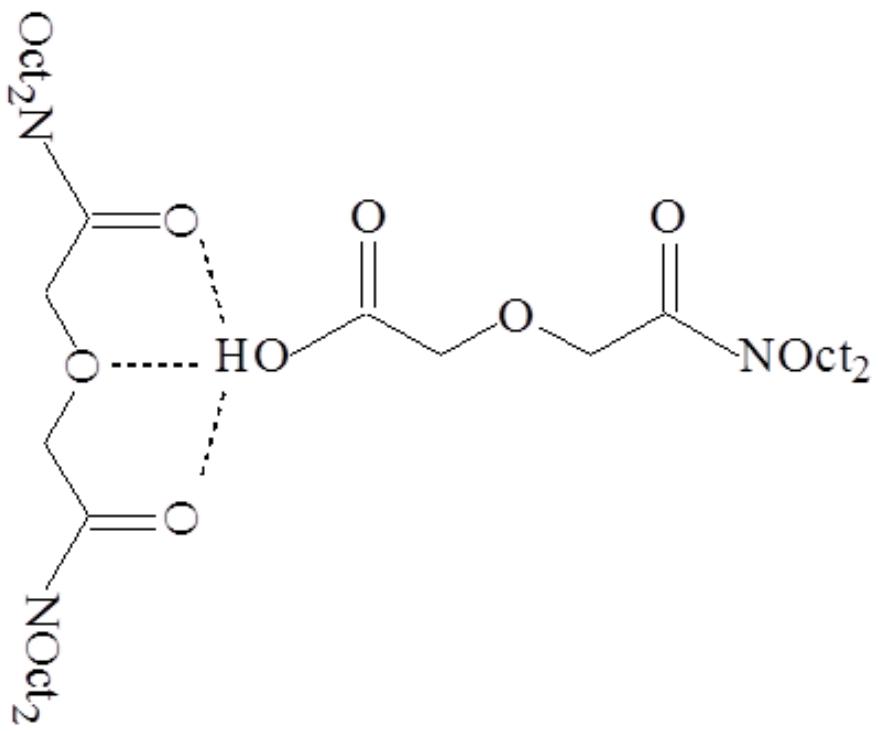
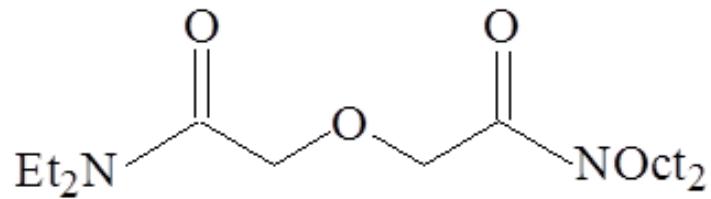
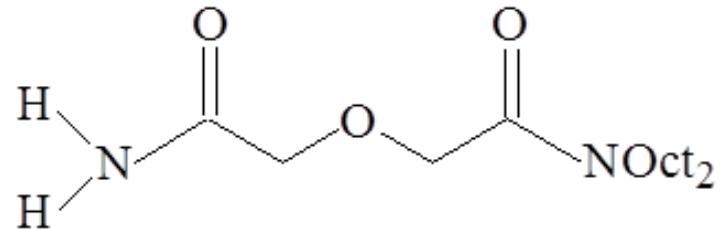
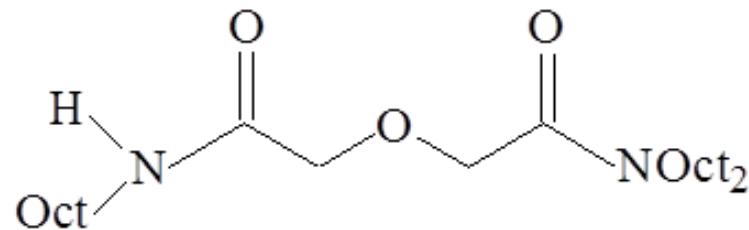
# Synthesis and Degradation of Dioctylamine



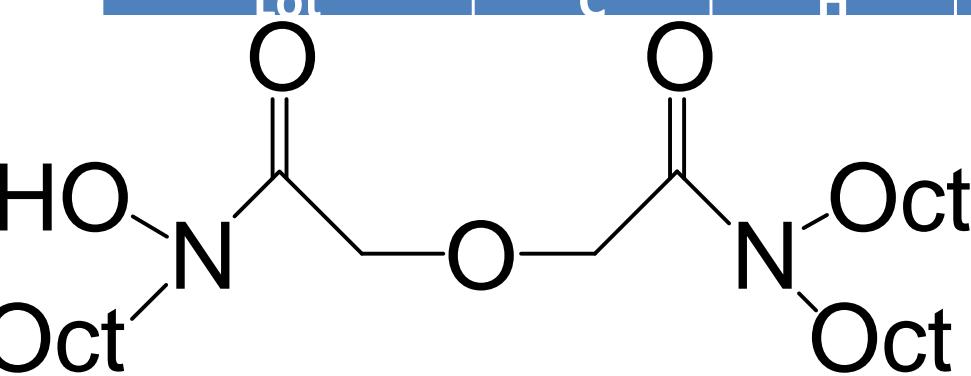
Octylamine impurity?

Hydrolysis/Oxidation by air/water?

# Common DGA Impurities by LC-MS



# Elemental Analysis of DGA Lots

Lot	(74.4%) C	(12.5%) H	(8.3%) O	(4.8%) N	S	P
 The chemical structure of Di(2-ethylhexyl) azodicarboxylate (DGA) is shown. It consists of two 2-ethylhexyl groups linked by an oxygen atom. Each group has a nitrogen atom bonded to it, which is further bonded to a hydroxyl group (HO) and another nitrogen atom. This second nitrogen atom is also bonded to an oxygen atom, which is part of a carbonyl group (C=O). The entire molecule is labeled "Oct" at both ends.	8.4	4.8	<0.07	<25ppm		
	C = 69.4%	O = 13.2%		N = 5.8%		
	H = 11.6%					
Impurity from Column	67.7	11.8	14.3	5.5	<0.05	<25ppm
DNX073115 (NEW DGC)	74.8	12.9	8.4	4.8	<0.05	<25ppm

# Summary

- DGA Synthesis and Purification Improved
- DGA QC updated with Po testing
- DGA lots from 2010-2015 tested retroactively
  - Standard Separations
  - Po Separations
    - Only separation affected by impurity
    - Other customer separations (as comments received)
    - Data available upon request